

Name: \_\_\_\_\_KEY\_\_\_\_\_

## Nomenclature D – Naming Acids

Please name the following acids.

(-ate becomes -ic and -ite becomes -ous for polyatomic anions ONLY!)

1. HBr (aq) hydrobromic acid
2. HCl (aq) hydrochloric acid
3. HI (aq) hydroiodic acid
4. HF (aq) hydrofluoric acid
5. H<sub>2</sub>S (aq) hydrosulfuric acid
6. HCN (aq) cyanic acid
7. HOH (l) water !!! (H<sub>2</sub>O)
8. HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub> (aq) acetic acid
9. H<sub>2</sub>CrO<sub>4</sub> (aq) chromic acid
10. H<sub>2</sub>CO<sub>3</sub> (aq) carbonic acid
11. HNO<sub>3</sub> (aq) nitric acid
12. H<sub>2</sub>SO<sub>4</sub> (aq) sulfuric acid
13. H<sub>3</sub>PO<sub>4</sub> (aq) phosphoric acid
14. HNO<sub>2</sub> (aq) nitrous acid
15. H<sub>2</sub>SO<sub>3</sub> (aq) sulfurous acid

Please indicate write the proper formula for the following acids.

Name	Cation	Anion	Formula
1. sulfuric acid	$\text{H}^+$	$\text{SO}_4^{2-}$	$\text{H}_2\text{SO}_4$ (aq)
2. phosphoric acid	$\text{H}^+$	$\text{PO}_4^{3-}$	$\text{H}_3\text{PO}_4$ (aq)
3. hydrobromic acid	$\text{H}^+$	$\text{Br}^-$	$\text{HBr}$ (aq)
4. cyanic acid	$\text{H}^+$	$\text{CN}^-$	$\text{HCN}$ (aq)
5. nitrous acid	$\text{H}^+$	$\text{NO}_2^-$	$\text{HNO}_2$ (aq)
6. carbonic acid	$\text{H}^+$	$\text{CO}_3^{2-}$	$\text{H}_2\text{CO}_3$ (aq)
7. acetic acid	$\text{H}^+$	$\text{C}_2\text{H}_3\text{O}_2^-$	$\text{HC}_2\text{H}_3\text{O}_2$ (aq)
8. chromic acid	$\text{H}^+$	$\text{CrO}_4^{2-}$	$\text{H}_2\text{CrO}_4$ (aq)
9. hydrosulfuric acid	$\text{H}^+$	$\text{S}^{2-}$	$\text{H}_2\text{S}$ (aq)
10. hydroiodic acid	$\text{H}^+$	$\text{I}^-$	$\text{HI}$ (aq)
11. nitric acid	$\text{H}^+$	$\text{NO}_3^-$	$\text{HNO}_3$ (aq)
12. hydrochloric acid	$\text{H}^+$	$\text{Cl}^-$	$\text{HCl}$ (aq)
13. hydrofluoric acid	$\text{H}^+$	$\text{F}^-$	$\text{HF}$ (aq)
14. sulfurous acid	$\text{H}^+$	$\text{SO}_3^{2-}$	$\text{H}_2\text{SO}_3$ (aq)