

Name: _____

Nomenclature B– Naming Transition Ions

Use the stock system to name the cation and add –ide to the end of the anion.

Give the names of the following compounds:

1. AgCl _____

2. FeI₃ _____

3. MnO _____

4. ZnF₂ _____

5. FeN _____

6. CuI _____

7. Cr₂S₃ _____

8. Cd₃P₂ _____

9. HgBr₂ _____

10. CuO _____

11. Mn₃N₄ _____

12. CoF₂ _____

13. SnO₂ _____

14. PbS _____

15. Hg₂I₂ _____

16. Co₃P₄ _____

17. SnBr₂ _____

18. PbS₂ _____

19. Ag₂O _____

20. CoCl₃ _____

Give the formulas of the following compounds:

<u>Name</u>	<u>Cation</u>	<u>Anion</u>	<u>Compound</u>
1. copper (I) fluoride	_____	_____	_____
2. tin (II) sulfide	_____	_____	_____
3. cobalt (IV) nitride	_____	_____	_____
4. iron (III) oxide	_____	_____	_____
5. mercury (I) bromide	_____	_____	_____
6. manganese (IV) sulfide	_____	_____	_____
7. lead (II) iodide	_____	_____	_____
8. iron (II) phosphide	_____	_____	_____
9. copper (II) oxide	_____	_____	_____
10. tin (IV) chloride	_____	_____	_____
11. silver sulfide	_____	_____	_____
12. mercury (II) nitride	_____	_____	_____
13. lead (IV) oxide	_____	_____	_____
14. cobalt (II) fluoride	_____	_____	_____
15. zinc bromide	_____	_____	_____