

Please memorize these ions for the next Exam
(*in addition* to the ions elements 1-20 form)

1. Co^{2+} = cobalt (II) ion
2. Co^{3+} = cobalt (III) ion
3. Co^{4+} = cobalt (IV) ion
4. Cu^+ = copper (I) ion
5. Cu^{2+} = copper (II) ion
6. Pb^{2+} = lead (II) ion
7. Pb^{4+} = lead (IV) ion
8. Fe^{2+} = iron (II) ion
9. Fe^{3+} = iron (III) ion
10. Sn^{2+} = tin (II) ion
11. Sn^{4+} = tin (IV) ion
12. Hg_2^{2+} = mercury (I) ion (actually two Hg^{+1} ions bonded together)
13. Hg^{2+} = mercury (II) ion
14. Mn^{2+} = manganese (II) ion
15. Mn^{4+} = manganese (IV) ion
16. Cr^{3+} = chromium (III) ion
17. Cd^{2+} = cadmium ion (some books use Roman and some don't, we will NOT)
18. Sr^{2+} = strontium ion (you *do not* need the Roman numeral for this ion)
19. Ni^{2+} = nickel ion (you *do not* need the Roman numeral for this ion)
20. Zn^{2+} = zinc ion (you *do not* need the Roman numeral for this ion)
21. Ag^+ = silver ion (you *do not* need the Roman numeral for this ion)