Net Ionic Equations Worksheet KEY

1. $Zn\ (s) + \mathbf{2}\ Ag^{+}\ (aq) \Rightarrow Zn^{2+}\ (aq) + \mathbf{2}\ Ag\ (s)$

2. $\mathbf{2}\ Na\ (s) + \mathbf{2}\ H_2O\ (l) \Rightarrow \mathbf{2}\ Na^{+}\ (aq) + \mathbf{2}\ OH^{-}\ (aq) + H_2\ (g)$

3. $Cl^{-}\ (aq) + Ag^{+}\ (aq) \Rightarrow AgCl\ (s)$

4. $\mathbf{3}\ Ca^{2+}\ (aq) + \mathbf{6}\ OH^{-}\ (aq) + \mathbf{2}\ H_3PO_4\ (aq) \Rightarrow \mathbf{6}\ H_2O\ (l) + Ca_3(PO_4)_2\ (s)$

5. $\mathbf{2}\ H^{+}\ (aq) + Ni\ (s) \Rightarrow Ni^{+2}\ (aq) + H_2\ (g)$

6. $Ba(HCO_3)_2\ (s) + \mathbf{2}\ H^{+}\ (aq) \Rightarrow Ba^{+2}\ (aq) + \mathbf{2}\ H_2O\ (g) + \mathbf{2}\ CO_2\ (g)$

7. $Ba^{+2}\ (aq) + SO_4^{-2}\ (aq) \Rightarrow BaSO_4\ (s)$

8. $Al_2(CO_3)_3\ (s) + 6\ H^{+}\ (aq) \Rightarrow 2\ Al^{+3}\ (aq) + 3\ H_2O\ (l) + 3\ CO_2\ (g)$

9. $Ca\ (s) + \mathbf{2}\ H_2O\ (l) \Rightarrow Ca^{2+}\ (aq) + \mathbf{2}\ OH^{-}\ (aq) + H_2\ (g)$

10. $Ag^{+}\ (aq) + CO_3^{-2}\ (aq) \Rightarrow Ag_2CO_3\ (s)$

11. $NaHCO_3\ (s) + H^{+}\ (aq) \Rightarrow Na^{+}\ (aq) + H_2O\ (l) + CO_2\ (g)$

12. $Mg^{+2}\ (aq) + \mathbf{2}\ OH^{-}\ (aq) \Rightarrow Mg(OH)_2\ (s)$

13. $Mn\ (s) + \mathbf{2}\ Cu^{+}\ (aq) \Rightarrow \mathbf{2}\ Cu\ (s) + Mn^{+2}\ (aq)$

14. $Zn\ (s) + \mathbf{2}\ H^{+}\ (aq) \Rightarrow Zn^{+2}\ (aq) + H_2\ (g)$

15. $\mathbf{3}\ OH^{-}\ (aq) + H_3PO_4\ (aq) \Rightarrow \mathbf{3}\ H_2O\ (l) + PO_4^{3-}\ (aq)$

16. $\mathbf{2}\ K\ (s) + \mathbf{2}\ H_2O\ (l) \Rightarrow \mathbf{2}\ K^{+}\ (aq) + \mathbf{2}\ OH\ (aq) + H_2\ (g)$

17. $Al^{+3}\ (aq) + PO_4^{+3}(aq) \Rightarrow AlPO_4\ (s)$

18. $\mathbf{2}\ OH^{-}\ (aq) + H_2CO_3\ (aq) \Rightarrow \mathbf{2}\ H_2O\ (l) + CO_3^{2-}\ (aq)$

19. $Cd\ (s) + Ni^{+2}\ (aq) \Rightarrow Cd^{+2}(aq) + Ni\ (s)$